

Davis - 7th Grade Science Agenda

Week of December 12, 2016

Day	In Class/Learning Targets	HW/Reminders
Monday 12-12	Snow Day	
<p style="text-align: center;">Tuesday 12-13</p> <p><i>I can understand how the total number of atoms does not change in a chemical reaction and the mass is conserved.</i></p>	<p style="text-align: center;">Block Schedule-Odd Day (3, 5)</p> <p>Finish and Check: Mystery Powders Lab</p> <ol style="list-style-type: none"> 1. Law of Conservation of Mass Notes and Practice 2. Law of Conservation of Mass Lab 3. Study Guide 	<p style="text-align: center;">Unit Test Dec. 19/20</p> <p><u>Success Criteria</u> *Identify the mystery powders</p> <p>* Totaling up the total amount of mass before and after a chemical reaction.</p>
Wednesday 12-14	Block Schedule-Even Day (2,4) See above	
<p style="text-align: center;">Thursday 12-15</p> <p><i>I can understand how the total number of atoms does not change in a chemical reaction and the mass is conserved.</i></p>	<p style="text-align: center;">See All Classes</p> <p>Check: Law of Conservation of Mass Lab</p> <ol style="list-style-type: none"> 1. Conservation of Mass Practice 2. Start Holiday Crystals Lab 	<p style="text-align: center;">Unit Test Dec. 19/20</p> <p><u>Success Criteria</u> * Correctly finding the total mass before and after a chemical reaction.</p> <p>Finish Study Guide!</p>
<p style="text-align: center;">Friday 12-16</p> <p><i>I can understand how the total number of atoms does not change in a chemical reaction and the mass is conserved.</i></p>	<p style="text-align: center;">See All Classes-Early Release</p> <ol style="list-style-type: none"> 1. Conservation of Mass Practice cont. 2. Check/Review Study Guide 	<p style="text-align: center;">Unit Test Dec. 19/20</p> <p><u>Success Criteria:</u> At least 20/25 on the study guide.</p>

Standards Covered This Unit:

MS-PS1-1 Develop models to describe the atomic composition of simple molecules and extended structures.

PS1.A: Disciplinary Core Ideas

- Substances are made from different types of atoms, which combine with one another in various ways. Atoms form molecules that range in size from two to thousands of atoms.
- Gases and liquids are made of molecules or inert atoms that are moving relative to each other.
- In a liquid, the molecules are constantly in contact with others; in a gas, they are widely spaced except when they happen to collide. In a solid, atoms are closely spaced and may vibrate in position but do not change relative locations
- Solids may be formed from molecules, or they may be extended structures with repeating subunits (e.g., crystals).
- The changes of state that occur with variations in temperature or pressure can be described and predicted using these models of matter.

MS-PS.2 Analyze and interpret data on the properties of substances before and after the substances interact to determine if a chemical reaction has occurred.

PS1.A: Disciplinary Core Ideas

Structure and Properties of Matter

- Each pure substance has characteristic physical and chemical properties (for any bulk quantity under given conditions) that can be used to identify it.

Patterns

- Macroscopic patterns are related to the nature of microscopic and atomic-level structure.